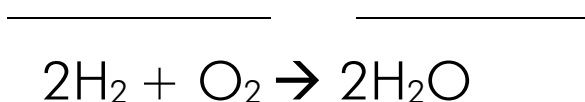




## Chemical Reactions

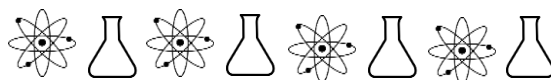
Fill in the blanks as you watch the video.

1. A chemical reaction is the process of one or more substances \_\_\_\_\_ to form new substances with different properties.
2. In chemical reactions, a new substance is formed from chemicals \_\_\_\_\_ with each other.
3. \_\_\_\_\_ are substances that enter a chemical reaction, while \_\_\_\_\_ are substances that are produced by a chemical reaction.
4. A chemical \_\_\_\_\_ is an expression using chemical symbols to represent a chemical reaction.
5. A plus sign is used to show that substances \_\_\_\_\_.
6. An \_\_\_\_\_ is used to show products yielded by reactants.
7. Label the reactants and the products in this chemical equation:



8. The law of the \_\_\_\_\_ of mass says that mass cannot be gained or lost in a chemical reaction.
9. The number of \_\_\_\_\_ of each element must be the same before and after a chemical reaction.
10. A \_\_\_\_\_ reaction is where two or more simple substances combine to form a more complex substance.

(continued on next page)



## Chemical Reactions cont.

11. A \_\_\_\_\_ reaction is where a substance breaks down into two or more simple substances.
12. A \_\_\_\_\_ - \_\_\_\_\_ reaction is where atoms of one element replace atoms of another element in a compound.
13. A \_\_\_\_\_ - \_\_\_\_\_ reaction is where atoms in two different compounds trade places with each other.
14. Chemical reactions involve \_\_\_\_\_ being given off or being absorbed.
15. An \_\_\_\_\_ reaction releases energy and gives off heat.
16. An \_\_\_\_\_ reaction absorbs energy resulting in the lowering of temperature.
17. Variables such as temperature, surface area, and concentration affect the \_\_\_\_\_ of chemical reactions, or the speed with which reactants turn into products.
18. The \_\_\_\_\_ \_\_\_\_\_ is the amount of material that comes in contact with other reactants.
19. \_\_\_\_\_ is the amount of substance in a given unit of volume.
20. A \_\_\_\_\_ is a substance that increases the reaction rate but is not changed by the reaction.